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# NEET - 2013 <br> (Physics, Chemistry and Biology) Code W 

## General Instructions:

1. The Answer sheet is inside this Text booklet. When you are directed to open the text booklet, take out the Answer Sheet and fill in the particulars on side - 1 and side - 2 carefully with blue/black ball point pen only.
2. The test is of 3 hours duration and consists of 180 questions. Each question carries 4 marks. For each correct response the candidate will get 4 marks. For each incorrect response, one mark will be deducted. The maximum marks are 720.
3. Use Blue / Black ball point pen only for writing particulars on this page / marking responses.
4. Rough work is to be done on the space provided for this purpose in the text booklet only.
5. On completion of the test, the candidate must handover the answer sheet to the invigilator in the room/ Hall. The candidates are allowed to take away this text booklet with them.
6. Make sure that the CODE printed on side - 2 of the answer sheet is the same as that on this booklet, In case of discrepancy, the candidate should immediately report the matter to the invigilator for the replacement of both the test Booklet and the Answer Sheet.
7. The candidates should ensure that the Answer sheet is not folded. Do not make any stray marks on the Answer sheet. Do not write your roll no. anywhere else except in the specified space in the Test booklet / Answer Sheet.
8. Use of white fluid for correction is not permissible on the Answer Sheet.

## Physics

1. In an experiment four quantities $a, b, c$ and $d$ are measured with percentage error $1 \%$, $2 \%, 3 \%$ and $4 \%$ respectively. Quantity $P$ is calculated as follows: $P=\frac{a_{2}}{b_{2}}$. \% error in P is
(1) $14 \%$
(2) $10 \%$
(3) $7 \%$
(4) $4 \%$
2. The velocity of a projectile at the initial point $A$ is $(2 \hat{i}+3 \hat{j}) \mathrm{m} / \mathrm{s}$. Its velocity (in $\mathrm{m} / \mathrm{s})$ at point $B$ is

(1) $-2 \hat{i}-3 \hat{j}$
(2) $-2 \hat{i}+3 \hat{j}$
(3) $2 \hat{i}-3 \hat{j}$
(4) $2 \hat{i}+3 \hat{j}$
3. A stone falls freely under gravity. It covers distances $h_{1}, h_{2}$ and $h_{3}$ in the first 5 seconds, the next 5 seconds and the next 5 seconds respectively. The relation between $h_{1}, h_{2}$ and $h_{3}$ is
(1) $\mathrm{h}_{1}=2 \mathrm{~h}_{2}=3 \mathrm{~h}_{2}$
(2) $\mathrm{h}_{1}=\underline{h}_{3^{1}}=\underline{h}_{5^{3}}$
(3) $h_{2}=3 h_{1}$ and $h_{3}=3 h_{2}$
(4) $h_{1}=h_{2}=h_{3}$
4. Three blocks with masses $\mathrm{m}, 2 \mathrm{~m}$ and 3 m are connected by strings, as shown in the figure. After an upward force $F$ is applied on block $m$, the masses move upward at constant speed v . What is the net force on the block of mass 2 m ? ( g is the acceleration due to gravity)

(1) Zero
(2) 2 mg
(3) 3 mg
(4) 6 mg
5. The upper half of an inclined plane of inclination $\theta$ is perfectly smooth while lower half is rough. A block starting from rest at the top of the plane will again come to rest at the bottom, if the coefficient of friction between the block and lower half of the plane is given by
(1) $\mu=\tan ^{1} \theta$
(2) $\mu=\tan ^{2} \theta$
(3) $\mu=2 \tan \theta$
(4) $\mu=\tan \theta$
6. A uniform force of $(3 \hat{i}+\hat{j})$ newton acts on a particle of mass 2 kg . Hence the particle is displaced from position $(2 \hat{i}+\hat{j})$ metre to position $(4 \hat{i}+3 \hat{j}-k)$ metre. The work done by the force on the particle is
(1) 9 J
(2) 6 J
(3) 13 J
(4) 15 J
7. An explosion breaks a rock into three parts in a horizontal plane. Two of them go off at right angles to each other. The first part of mass 1 kg moves with a speed of $12 \mathrm{~ms}^{-}$ 1 and the second part of mass 2 kg moves with $8 \mathrm{~ms}^{-1}$ speed. If the third part flies off with $4 \mathrm{~ms}^{-1}$ speed, then its mass is
(1) 3 kg
(2) 5 kg
(3) 7 kg
(4) 17 kg
8. A rod $P Q$ of mass $M$ and length $L$ is hinged at end $P$. The rod is kept horizontal by a massless string tied to point Q as shown in figure. When string is cut, the initial angular acceleration of the rod is

(1) ${ }^{3 g} 2 \mathrm{~L}$
(2) $\mathrm{L}^{\mathrm{g}}$
(3) $\frac{2 g_{L}}{}$
(4) $3 \mathrm{~L}^{2 \mathrm{~g}}$
9. A small object of uniform density rolls up a curved surface with an initial velocity ' $v$ '. It reaches up to a maximum height of $\frac{3 v_{2}}{4 g}$ with respect to the initial position.
The object is
(1) Ring
(2) Solid sphere
(3) Hollow sphere
(4) Disc
10.A body of mass ' m ' taken from the earth's surface to the height equal to twice the radius $(R)$ of the earth. The change in potential energy of body will be
(1) Mg 2 R
(2) $3^{2} \mathrm{mgR}$
(3) 3 mgR
(4) $\underline{1}_{3 \mathrm{mgR}}$
10. Infinite number of bodies, each of mass 2 kg is situated on x -axis at distance $1 \mathrm{~m}, 2 \mathrm{~m}$, $4 \mathrm{~m}, 8 \mathrm{~m}$, respectively, from the origin. The resulting gravitational potential due to this system at the origin will be
(1) -G
(2) $-\underline{8}_{3 G}$
(3) $-3^{4_{G}}$
(4) -4 G
12.The following four wires are made of the same material. Which of these will have the largest extension when the same tension is applied?
(1) Length $=50 \mathrm{~cm}$, diameter $=0.5 \mathrm{~mm}$
(2) Length $=100 \mathrm{~cm}$, diameter $=1 \mathrm{~mm}$
(3) Length $=200 \mathrm{~cm}$, diameter $=2 \mathrm{~mm}$
(4) Length $=300 \mathrm{~cm}$, diameter $=3 \mathrm{~mm}$
11. The wettability of a surface by a liquid depends primarily on
(1) Viscosity
(2) Surface tension
(3) Density
(4) Angle of contact between the surface and the liquid
12. The molar specific heats of an ideal gas at constant pressure and volume are denoted by $C_{p}$ and $C_{v}$ respectively. If $\gamma=\frac{C_{p}}{C_{v}}$ and $R$ is the universal gas constant, then $C \quad{ }_{v}$ is equal to
(1) $1+\gamma$
$1-\gamma$
(2) $\frac{\mathrm{R}}{(\gamma-1)}$
(3) $\frac{(\gamma-1)}{R}$
(4) $\gamma \mathrm{R}$
13. A piece of iron is heated in a flame. It first becomes dull red then becomes reddish yellow and finally turns to white hot. The correct explanation for the above observation is possible by using
(1) Stefan's Law
(2) Wien's displacement Law
(3) Kirchoff's Law
(4) Newton's Law of cooling
14. A gas is taken through the cycle $A \rightarrow B \rightarrow C \rightarrow A$, as shown. What is the net work done by the gas?

(1) 2000 J
(2) 1000 J
(3) Zero
(4) -2000 J
15. During an adiabatic process, the pressure of a gas is found to be proportional to the cube of its temperature. The ratio of $C_{p \text { for the gas is }} C_{v}$
(1) $3^{4}$
(2) 2
(3) $\underline{5}_{3}$
(4) $\underline{3}_{2}$
16. In the given $(\mathrm{V}-\mathrm{T})$ diagram, what is the relation between pressures $\mathrm{P}_{1}$ and $\mathrm{P}_{2}$ ?

(1) $P_{2}=P_{1}$
(2) $\mathrm{P}_{2}>\mathrm{P}_{1}$
(3) $P_{2}<P_{1}$
(4) Cannot be predicted
17. The amount of heat energy required to raise the temperature of 1 g of Helium at NTP, from $\mathrm{T}_{1} \mathrm{~K}$ to $\mathrm{T}_{2} \mathrm{~K}$ is

3
(1) $8 \mathrm{~N}_{\mathrm{a}} \mathrm{k}_{\mathrm{B}}\left(\mathrm{T}_{2}-\mathrm{T}_{1}\right)$
$\underline{3}$
(2) $2 \mathrm{~N}_{\mathrm{a}} \mathrm{k}_{\mathrm{B}}\left(\mathrm{T}_{2}-\mathrm{T}_{1}\right)$
$\underline{3}$
(3) $4 \mathrm{Na}_{\mathrm{a}} \mathrm{k}$ в $\left(\mathrm{T}_{2}-\mathrm{T}_{1}\right)$
(4) $\frac{3}{4} N_{a}\left(\frac{T}{\mathrm{~K}_{B} \mid} \left\lvert\, \begin{array}{l}T^{2} \mid \\ 1 \\ 1\end{array}\right.\right)$
20.A wave travelling in the +ve x -direction having displacement along y -direction as 1 m , 1
wavelength $2 \pi \mathrm{~m}$ and frequency of $\pi \mathrm{Hz}$ is represented by
(1) $Y=\sin (x-2 t)$
(2) $Y=\sin (2 \pi x-2 \pi t)$
(3) $Y=\sin (10 \pi x-20 \pi t)$
(4) $Y=\sin (2 \pi x-2 \pi t)$
21. If we study the vibration of a pipe open at both ends, then the following statement is not true
(1) Open end will be anti-node
(2) Odd harmonics of the fundamental frequency will be generated
(3) All harmonics of the fundamental frequency will be generated
(4) Pressure change will be maximum at both ends
22.A source of unknown frequency gives 4 beats/s, when sounded with a source of known frequency 250 Hz . The second harmonic of the source of unknown frequency gives five beats per second, when sounded with a source of frequency 513 Hz . The unknown frequency is
(1) 254 Hz
(2) 246 Hz
(3) 240 Hz
(4) 260 Hz
23. Two pith balls carrying equal charges are suspended from a common point by strings of equal length; the equilibrium separation between them is r. Now the strings are rigidly clamped at half the height. The equilibrium separation between the balls now become

24. $A, B$ and $C$ are three points in a uniform electric field. The electric potential is

(1) Maximum at A
(2) Maximum at B
(3) Maximum at C
(4) Same at all the three points A, B and C
25. A wire of resistance $4 \Omega$ is stretched to twice its original length. The resistance of stretched wire would be
(1) $2 \Omega$
(2) $4 \Omega$
(3) $8 \Omega$
(4) $16 \Omega$
26. The internal resistance of a 2.1 V cell which gives a current of 0.2 A through a resistance of $10 \Omega$ is
(a) $0.2 \Omega$
(b) $0.5 \Omega$
(c) $0.8 \Omega$
(d) $1.0 \Omega$
27. The resistances of the four arms $P, Q, R$ and $S$ in a Wheatstone's bridge are $10 \mathrm{ohm}, 30$ ohm, 30 ohm and 90 ohm, respectively. The e.m.f. and internal resistance of the cell are 7 volt and 5 ohm respectively. If the galvanometer resistance is 50 ohm, the current drawn from the cell will be
(1) 1.0 A
(2) 0.2 A
(3) 0.1 A
(4) 2.0 A
28. When a proton is released from rest in a room, it starts with an initial acceleration a0 towards west. When it is projected towards north with a speed vo it moves with an initial acceleration 3ao toward west. The electric and magnetic fields in the room are
(1) $\xrightarrow{\text { ma }_{0}}$ west,,$\underline{2 \mathrm{ma}_{0}}$
up evo
(2) $\frac{\text { ma }}{e} 0$ west, $\frac{2 m a 0}{{\underset{e}{0}}_{0}^{e}}$ down
(3) $\frac{m a_{0}}{e}$ east, $\frac{3 \text { mao }}{e v}$ up
(4) $\frac{\mathrm{ma}_{0}}{\mathrm{e}}$ east, $\frac{3 m a_{0}}{\text { ev }}{ }_{0}$ down
29. A current loop in a magnetic field
(1) Experiences a torque whether the field is uniform or non uniform in all orientations
(2) Can be in equilibrium in one orientation
(3) Can be in equilibrium in two orientations, both the equilibrium states are unstable
(4) Can be in equilibrium in two orientations, one stable while the other is unstable
30.A bar magnet of length $l$ and magnetic dipole moment $M$ is bent in the form of an arc as shown in figure. The new magnetic dipole moment will be

(1) M
(2) $\pi^{3} M$
(3) $\bar{\pi}^{2} \mathrm{M}$
(4) $\underline{M}_{2}$
31. A wire loop is rotated in a magnetic field. The frequency of change of direction of the induced e.m.f. is
(1) Once per revolution
(2) Twice per revolution
(3) Four times per revolution
(4) Six times per revolution
32.A coil of self-inductance $L$ is connected in series with a bulb $B$ and an AC source. Brightness of the bulb decreases when
(1) Frequency of the AC source is decreased
(2) Number of turns in the coil is reduced
(3) A capacitance of reactance $X_{C}=X_{L}$ is included in the same circuit
(4) An iron rod is inserted in the coil
33. The condition under which a microwave oven heats up a food item containing water molecules most efficiently is
(1) The frequency of the microwaves must match the resonant frequency of the water molecules
(2) The frequency of the microwaves has no relation with natural frequency of water molecules.
(3) Microwaves are heat waves, so always produce heating
(4) Infra-red waves produce heating in a microwave oven
34. Ratio of longest wavelengths corresponding to Lyman and Balmer series in hydrogen spectrum is
(1) $27^{5}$
(2) $23^{3}$
(3) $29^{7}$
(4) $31^{9}$
35. The half life of a radioactive isotope ' X ' is 20 years. It decays to another element ' Y ' which is stable. The two elements ' X ' and ' Y ' were found to be in the ratio $1: 7$ in a sample of a given rock. The age of the rock is estimated to be
(1) 40 years
(2) 60 years
(3) 80 years
(4) 100 years
36. A certain mass of Hydrogen is changed to Helium by the process of fusion. The mass defect in fusion reaction is 0.02866 u . The energy liberated per $u$ is
(given $1 \mathrm{u}=931 \mathrm{MeV}$ )
(1) 2.67 MeV
(2) 26.7 MeV
(3) 6.675 MeV
(4) 13.35 MeV
37. For photoelectric emission from certain metal the cut-off frequency is $v$. If radiation of frequency $2 v$ impinges on the metal plate, the maximum possible velocity of the emitted electron will be ( m is the electron mass)
(1) $\sqrt{\frac{h \nu}{(2 m)}}$
(2) $\sqrt{\frac{h v}{m}}$
(3) $\sqrt{\frac{2 h v}{m}}$
(4) $2 \sqrt{h_{m}^{m}} v$
38. The wavelength $\lambda_{e}$ of an electron and $\lambda_{p}$ of a photon of same energy $E$ are related by
(1) $\lambda_{p} \propto \lambda^{2} e$
(2) $\lambda_{p} \propto \lambda_{e}$
(3) $\lambda_{p} \propto \sqrt{\lambda_{e}}$
(4) $\lambda_{\mathrm{p}} \propto \frac{V_{1}}{\sqrt{ } \lambda_{e}}$
39. A plano-convex lens fits exactly into a plano concave lens. Their plane surfaces are parallel to each other. If lenses are made of different materials of refractive indices $\mu_{1}$ and $\mu_{2}$ and R is the radius of curvature of the curved surface of the lenses, then the focal length of the combination is
(1) $\frac{R}{2\left(\mu_{1}+\mu_{2}\right)}$

R
(2) $2\left(\mu_{1}-\mu_{2}\right)$

R
(3) $\left(\mu_{1}-\mu_{2}\right)$
(4) $\frac{2 R}{\left(\mu_{2}-\mu_{1}\right)}$
40.For a normal eye, the cornea of eye provides a converging power of 40 D and the least converging power of the eye lens behind the cornea is 20 D . Using this information, the distance between the retina and the cornea - eye lens can be estimated to be
(1) 5 cm
(2) 2.5 cm
(3) 1.67 cm
(4) 1.5 cm
41.In Young's double slit experiment, the slits are 2 mm apart and are illuminated by photons of two wavelengths $\lambda_{1}=12000 \AA$ and $\lambda_{2}=10000 \AA$. At what minimum distance from the common central bright fringe on the screen 2 m from the slit will a bright fringe from one interference pattern coincide with a bright fringe from the other?
(1) 8 mm
(2) 6 mm
(3) 4 mm
(4) 3 mm
42.A parallel beam of fast moving electrons is incident normally on a narrow slit. A fluorescent screen is placed at a large distance from the slit. If the speed of the electrons is increased, which of the following statements is correct?
(1) Diffraction pattern is not observed on the screen in the case of electrons
(2) The angular width of the central maximum of the diffraction pattern will increase
(3) The angular width of the central maximum will decrease
(4) The angular width of the central maximum will be unaffected
43. In a n-type semiconductor, which of the following statement is true?
(1) Electrons are majority carriers and trivalent atoms are dopants
(2) Electron are minority carriers and pentavalent atoms are dopants
(3) Holes are minority carriers and pentavalent atoms are dopants
(4) Holes are majority carriers and trivalent atoms are dopants
44. In a common emitter (CE) amplifier having a voltage gain G , the transistor used has transconductance 0.03 mho and current gain 25 . If the above transistor is replaced with another one with transconductance 0.02 mho and current gain 20 , the voltage gain will be
(1) $3^{\underline{Z_{G}}}$
(2) 1.5 G
(3) $\stackrel{1}{1}_{3 G}$
(4) $\underline{5}_{4 G}$
45. The output $(X)$ of the logic circuit shown in figure will be

(1) $X=\overline{\bar{A}} \overline{\bar{B}}$
(2) $X=\overline{A B}$
(3) $X=A \cdot B$
(4) $X=\overline{A+B}$

## Chemistry

46. The value of Planck's constant is $6.63 \times 10^{-34} \mathrm{Js}$. The speed of light is $3 \times 10^{17} \mathrm{~nm} \mathrm{~s}-1$. Which value is closest to the wavelength in nanometer of a quantum of light with frequency of $6 \times 10^{15} \mathrm{~s}^{-1}$ ?
(1) 10
(2) 25
(3) 50
(4) 75
47. What is the maximum numbers of electrons that can be associated with the following set of quantum numbers?
$\mathrm{n}=3, \mathrm{l}=1$ and $\mathrm{m}=-1$
(1) 10
(2) 6
(3) 4
(4) 2
48. What is the activation energy for a reaction if its rate doubles when the temperature is raised from $20^{\circ} \mathrm{C}$ to $35^{\circ} \mathrm{C} ?(\mathrm{R}=8.314 \mathrm{~J}$ mol $-1 \mathrm{~K}-1)$
(1) $342 \mathrm{~kJ} \mathrm{~mol}^{-1}$
(2) $269 \mathrm{~kJ} \mathrm{~mol}^{-1}$
(3) $34.7 \mathrm{~kJ} \mathrm{~mol}^{-1}$
(4) $15.1 \mathrm{~kJ} \mathrm{~mol}^{-1}$
49. A hydrogen gas electrode is made by dipping platinum wire in a solution of HCl of pH $=10$ and by passing hydrogen gas around the platinum wire at one atm pressure. The oxidation potential of electrode would be
(1) 0.059 V
(2) 0.59 V
(3) 0.118 V
(4) 1.18 V
50.A reaction having equal energies of activation for forward and reverse reactions has
(1) $\Delta S=0$
(2) $\Delta G=0$
(3) $\Delta H=0$
(4) $\Delta H=\Delta G=\Delta S=0$
50. At $25^{\circ} \mathrm{C}$ molar conductance of 0.1 molar aqueous solution of ammonium hydroxide is $9.54 \mathrm{ohm}^{-1} \mathrm{~cm}^{2} \mathrm{~mol}^{-1}$ and at infinite dilution its molar conductance is $238 \mathrm{ohm}^{-1}$ $\mathrm{cm}^{2} \mathrm{~mol}^{-1}$. The degree of ionisation of ammonium hydroxide at the same concentration and temperature is
(1) $2.080 \%$
(2) $20.800 \%$
(3) $4.008 \%$
(4) $40.800 \%$
51. Based on equation $\mathrm{E}=-2.178 \times 10^{-18} \mathrm{~J}\left(\underline{\mathrm{Z}_{2}}\right)$
$\left(\mathrm{n}^{2}\right)$${ }^{\text {certain }}$ conclusions are written. Which
of them is not correct?
(1) The negative sign in equation simply means that the energy of electron bound to the nucleus is lower than it would be if the electrons were at the infinite distance from the nucleus.
(2) Larger the value of $n$, the larger is the orbit radius.
(3) Equation can be used to calculate the change in energy when the electron changes orbit.
(4) For $n=1$, the electron has a more negative energy than it does for $n=6$ which means that the electron is more loosely bound in the smallest allowed orbit.
52. A button cell used in watches functions as following
$\mathrm{Zn}_{(\mathrm{s})}+\mathrm{Ag}_{2} \mathrm{O}_{(\mathrm{s})}+\mathrm{H}_{2} \mathrm{O}_{2} 2 \mathrm{Ag}_{(\mathrm{s})}+\mathrm{Zn}^{2+}{ }_{(\mathrm{aq})}+2 \mathrm{H}^{-}{ }_{(\mathrm{aq})}$
If half cellpotentialsare
$\mathrm{Zn}^{2+}(\mathrm{aq})+2 \mathrm{e}^{-} \rightarrow \mathrm{Zn}_{(\mathrm{s})} ; \mathrm{E}^{0}=-0.76 \mathrm{~V}$
$\mathrm{Ag}_{2} \mathrm{O}_{(\mathrm{s})}+\mathrm{H}_{2} \mathrm{O}_{(\mathrm{l})}+2 \mathrm{e}^{-} \rightarrow 2 \mathrm{Ag}_{(\mathrm{s})}+2 \mathrm{OH}^{-}(\mathrm{aq})$,
$\mathrm{E}^{0}=0.34 \mathrm{~V}$
The cell potential will be
(1) 1.10 V
(2) 0.42 V
(3) 0.84 V
(4) 1.34 V
53. How many grams of concentrated nitric acid solution should be used to prepare 250 mL of $2.0 \mathrm{M} \mathrm{HNO}_{3}$ ? The concentrated acid is $70 \% \mathrm{HNO}_{3}$.
(1) 45.0 g conc. $\mathrm{HNO}_{3}$
(2) 90.0 g conc. $\mathrm{HNO}_{3}$
(3) 70.0 g conc. $\mathrm{HNO}_{3}$
(4) 54.0 g conc. $\mathrm{HNO}_{3}$
54. The number of carbon atoms per unit cell of diamond unit cell is
(1) 4
(2) 8
(3) 6
(4) 1
55. Maximum deviation from ideal gas is expected from
(1) $\mathrm{H}_{2}(\mathrm{~g})$
(2) $\mathrm{N}_{2}(\mathrm{~g})$
(3) $\mathrm{CH}_{4}(\mathrm{~g})$
(4) $\mathrm{NH}_{3}(\mathrm{~g})$
56. A metal has a fcc lattice. The edge length of the unit cell is 404 pm . The density of the metal is $2.72 \mathrm{~g} \mathrm{~cm}^{-3}$. The molar mass of the metal
is ( $\mathrm{N}_{\mathrm{A}}$ Avogadro's constant $=6.02 \times 10^{23} \mathrm{~mol}^{-1}$ )
(1) $40 \mathrm{~g} \mathrm{~mol}^{-1}$
(2) $30 \mathrm{~g} \mathrm{~mol}^{-1}$
(3) $27 \mathrm{~g} \mathrm{~mol}^{-1}$
(4) $20 \mathrm{~g} \mathrm{~mol}^{-1}$
57. Dipole-induced dipole interactions are present in which of the following pairs?
(1) $\mathrm{H}_{2} \mathrm{O}$ and alcohol
(2) $\mathrm{Cl}_{2}$ and $\mathrm{CCl}_{4}$
(3) HCl and He atoms
(4) $\mathrm{SiF}_{4}$ and He atoms
58. A magnetic moment of 1.73 BM will be shown by one among the following
(1) $\left[\mathrm{Cu}\left(\mathrm{NH}_{3}\right)_{4}\right]^{2+}$
(2) $\left[\mathrm{Ni}(\mathrm{CN})_{4}\right]^{2-}$
(3) $\mathrm{TiCl}_{4}$
(4) $\left[\mathrm{CoCl}_{6}\right]^{4-}$
60.Roasting of sulphides gives the gas $X$ as a byproduct. This is a colorless gas with choking smell of burnt sulphur and causes great damage to respiratory organs as a result of acid rain. Its aqueous solution is acidic acts as a reducing agent and its acid has never been isolated. The gas X is
(1) $\mathrm{H}_{2} \mathrm{~S}$
(2) $\mathrm{SO}_{2}$
(3) $\mathrm{CO}_{2}$
(4) $\mathrm{SO}_{3}$
59. Which is the strongest acid in the following?
(1) $\mathrm{H}_{2} \mathrm{SO}_{4}$
(2) $\mathrm{HClO}_{3}$
(3) $\mathrm{HClO}_{4}$
(4) $\mathrm{H}_{2} \mathrm{SO}_{3}$
62.Which of the following is paramagnetic?
(1) CO
(2) $\mathrm{O}_{2}^{-}$
(3) $\mathrm{CN}^{-}$
(4) $\mathrm{NO}^{+}$
60. Which of the following structure is similar to graphite?
(1) BN
(2) B
(3) $\mathrm{B}_{4} \mathrm{C}$
(4) $\mathrm{B}_{2} \mathrm{H}_{6}$
64.The basic structural unit of silicates is
(1) $\mathrm{SiO}^{-}$
(2) $\mathrm{SiO}_{4}^{4}-$
(3) $\mathrm{SiO}_{3}{ }^{-}$
(4) $\mathrm{SiO}^{2}{ }^{-}$
61. Reaction by which Benzaldehyde cannot be prepared?
(1)
(2)




(4)

62. Which of the following does not give oxygen on heating?
(1) $\mathrm{KClO}_{3}$
(2) $\mathrm{Zn}\left(\mathrm{ClO}_{3}\right)_{2}$
(3) $\mathrm{K}_{2} \mathrm{Cr}_{2} \mathrm{O}_{7}$
(4) $\left(\mathrm{NH}_{4}\right)_{2} \mathrm{Cr}_{2} \mathrm{O}_{7}$
63. Which of the following lanthanoid ions is diamagnetic?
(At.no. $\mathrm{Ce}=58, \mathrm{Sm}=62, \mathrm{Eu}=63, \mathrm{Yb}=70$ )
(1) $\mathrm{Ce}^{2+}$
(2) $\mathrm{Sm}^{2+}$
(3) $\mathrm{Eu}^{2+}$
(4) $\mathrm{Yb}^{2+}$
64. Identify the correct order of solubility in aqueous medium
(1) $\mathrm{CuS}>\mathrm{ZnS}>\mathrm{Na} 2 \mathrm{~S}$
(2) $\mathrm{ZnS}>\mathrm{Na} 2 \mathrm{~S}>\mathrm{CuS}$
(3) $\mathrm{Na} 2 \mathrm{~S}>\mathrm{CuS}>\mathrm{ZnS}$
(4) $\mathrm{Na} 2 \mathrm{~S}>\mathrm{ZnS}>\mathrm{CuS}$
65. $\mathrm{XeF}_{2}$ is isostructural with
(1) $\mathrm{TeF}_{2}$
(2) $\mathrm{ICl}_{2}^{-}$
(3) $\mathrm{SbCl}_{3}$
(4) $\mathrm{BaCl}_{2}$
70.An excess of $\mathrm{AgNO}_{3}$ is added to 100 mL of a 0.01 M solution of dichlorotetraaquachromium(III) chloride. The number of moles of AgCl precipitated would be
(1) 0.001
(2) 0.002
(3) 0.003
(4) 0.01
66. Which of these is least likely to act as a Lewis base?
(1) CO
(2) $\mathrm{F}^{-}$
(3) $\mathrm{BF}_{3}$
(4) $\mathrm{PF}_{3}$
67. $\mathrm{KMnO}_{4}$ can be prepared from $\mathrm{K}_{2} \mathrm{MnO}_{4}$ as per the reaction:
$3 \mathrm{MnO}^{2}{ }_{4}^{-}+2 \mathrm{H}_{2} \mathrm{O} \rightarrow 2 \mathrm{MnO}_{4}^{-}+\mathrm{MnO}_{2}+4 \mathrm{OH}^{-}$
The reaction can go completion by removing $\mathrm{OH}^{-}$ions by adding:
(1) HCl
(2) KOH
(3) $\mathrm{CO}_{2}$
(4) $\mathrm{SO}_{2}$
68. Which of the following is electron-deficient?
(1) $\left(\mathrm{CH}_{3}\right)_{2}$
(2) $\left(\mathrm{SiH}_{3}\right)_{2}$
(3) $\left(\mathrm{BH}_{3}\right)_{2}$
(4) PH 3
74.Structure of the compound whose IUPAC name is 3- Ethyl-2-hydroxy-4-methylhex-3-en-5-ynoic acid is

(2)

(3)

(4)

69. Which of these is not a monomer for a high molecular mass silicone polymer?
(1) $\mathrm{MeSiCl}_{3}$
(2) $\mathrm{Me}_{2} \mathrm{SiCl}_{2}$
(3) Me 3 SiCl
(4) $\mathrm{PhSiCl}_{3}$
70. Which of the following statements about the interstitial compounds is incorrect?
(1) They retain metallic conductivity
(2) They are chemically reactive
(3) They are much harder than the pure metal
(4) They have higher melting points than the pure metal
71. Which one of the following molecules contains no $\pi$ bond?
(1) $\mathrm{CO}_{2}$
(2) $\mathrm{H}_{2} \mathrm{O}$
(3) $\mathrm{SO}_{2}$
(4) $\mathrm{NO}_{2}$
72. Antiseptics and disinfectants either kill or prevent growth of microorganisms. Identify which of the following statements is not true
(1) A $0.2 \%$ solution of phenol is an antiseptic while $1 \%$ solution acts as a disinfectant
(2) Chlorine and Iodine are used as strong disinfectants
(3) Dilute solutions of Boric acid and Hydrogen peroxide are strong antiseptics
(4) Disinfectants harm the living tissues
79.Among the following ethers, which one will produce methyl alcohol on treatment with hot concentrated HI ?
(1) $\mathrm{CH}_{3}-\mathrm{CH}_{2}-\mathrm{CH}_{2}-\mathrm{CH}_{2}-\mathrm{O}-\mathrm{CH}_{3}$
(2) $\mathrm{CH}_{3}-\mathrm{CH}_{2}-\mathrm{CH}-\mathrm{O}-\mathrm{CH}_{3}$
|
CH3
(3)

(4) $\mathrm{CH}_{3}-\mathrm{CH}-\mathrm{CH} 2-\mathrm{O}-\mathrm{CH}_{3}$

CH3
73. Nylon is an example of:
(1) Polyester
(2) Polysaccharide
(3) Polyamide
(4) Polythene
74. The structure of isobutyl group in an organic compound is:
(2)


$$
\mathrm{CH}_{3}-\mathrm{CH}-\mathrm{CH}_{2}-\mathrm{CH}_{3}
$$

(3) $\mathrm{CH}_{3}-\mathrm{CH}_{2}-\mathrm{CH}_{2}-\mathrm{CH}_{2}-$
(4)

82. Nitrobenzene on reaction with conc. $\mathrm{HNO}_{3} / \mathrm{H}_{2} \mathrm{SO}_{4}$ at $80-100^{\circ} \mathrm{C}$ forms which one of the following products?
(1) 1, 2-Dinitrobenzene
(2) 1, 3-Dinitrobenzene
(3) 1, 4-Dinitrobenzene
(4) 1, 2, 4-Trinitrobenzene
83. Some meta-directing substituents in aromatic substitution are given. Which one is most deactivating?
(1) $-\mathrm{C} \equiv \mathrm{N}$
(2) $-\mathrm{SO}_{3} \mathrm{H}$
(3) -COOH
(4) $-\mathrm{NO}_{2}$
$84.6 .02 \times 10^{20}$ molecules of urea are present in 100 mL of its solution. The concentration of solution is
(1) 0.02 M
(2) 0.01 M
(3) 0.001 M
(4) 0.1 M
85. Which of the following is a polar molecule?
(1) $\mathrm{BF}_{3}$
(2) $\mathrm{SF}_{4}$
(3) $\mathrm{SiF}_{4}$
(4) XeF4
86. Which is the monomer of Neoprene in the following?
(1) $\mathrm{CH}_{2}=\mathrm{CH}-\mathrm{C}=\mathrm{CH}_{2}$
(2) $\mathrm{CH}_{2}=\mathrm{C}-\mathrm{CH}=\mathrm{CH}_{2}$
|
$\mathrm{CH}_{3}$
(3) $\mathrm{CH}_{2}=\mathrm{C}-\mathrm{CH}=\mathrm{CH}_{2}$
|
Cl
(4) $\mathrm{CH}_{2}=\mathrm{CH}-\mathrm{CH} \equiv \mathrm{CH}_{2}$
87. In the reaction


A is
(1) $\mathrm{HgSO}_{4} / \mathrm{H}_{2} \mathrm{SO}_{4}$
(2) $\mathrm{Cu}_{2} \mathrm{Cl}_{2}$
(3) $\mathrm{H}_{3} \mathrm{PO}_{2}$ and $\mathrm{H}_{2} \mathrm{O}$
(4) $\mathrm{H}^{+} / \mathrm{H}_{2} \mathrm{O}$
88. The radical

(1) 6 p-orbitals and 6 unpaired electrons
(2) 7 p-orbitals and 6 unpaired electrons
(3) 7 p-orbitals and 7 unpaired electrons
(4) 6 p-orbitals and 7 unpaired electrons
89. The order of stability of the following tautomeric compounds is



II


III
(1) I $>$ II $>$ III
(2) III $>$ II $>$ I
(3) II $>$ I $>$ III
(4) II $>$ III $>$ I
90. Which of the following compounds will not undergo Friedal-Craft's reaction easily?
(1) Cumene
(2) Xylene
(3) Nitrobenzene
(4) Toluene

## Biology

91. Select the wrong statement:
(1) Isogametes are similar in structure, function and behaviour
(2) Anisogametes differ either in structure, function or behaviour
(3) In Oomycetes female gemete is smaller and motile, while male gamete is larger and nonmotile
(4) Chlamydomomas Exhibits both isogamy and anisogamy and Fucus shows oogamy
92. Which one of the following is not a correct statement?
(1) Herbarium houses dried, pressed and preserved plant specimens
(2) Botanical gardens have collection of living plants for reference.
(3) A museum has collection of photographs of plants and animals.
(4) Key is a taxonomic aid for identification of specimens.
93. Isogamous condition with non-flagellated gametes is found in
(1) Chlamydomonas
(2) Spirogyra
(3) Volvox
(4) Fucus
94. Besides paddy fields, cyanobacteria are also found inside vegetative part of
(1) Pinus
(2) Cycas
(3) Equisetum
(4) Psilotum
95. Megasporangium is equivalent to
(1) Embryo sac
(2) Fruit
(3) Nucellus
(4) Ovule
96. Read the following statements (A-E) and answer the question which follows them
(A) In liverworts, mosses, and ferns gametophytes are free-living
(B) Gymnosperms and some ferns are heterosporous
(C) Sexual reproduction in Fucus, Volvox and Albugo is oogamous
(D) The sporophyte in liverworts is more elaboratethan that in mosses
(E) Both, Pinus and Marchantia are dioecious

How many of the above statements are correct?
(1) One
(2) Two
(3) Three
(4) Four
97.Among bitter gourd, mustard, brinjal, pumpkin, china rose, lupin, cucumber, sunnhemp, gram, guava, bean, chilli, plum, petunia, tomato, rose, withania, potato, onion, aloe and tulip how many plants have hypogynous flower?
(1) Six
(2) Ten
(3) Fifteen
(4) Eighteen
98. Interfascicular cambium develops from the cells of
(1) Medullary rays
(2) Xylem parenchyma
(3) Endodermis
(4) Pericycle
99. In China rose the flowers are
(1) Actinomorphic, hypogynous with twisted aestivation
(2) Actinomorphic, epigynous with valvate aestivation
(3) Zygomorphic, hypogynous with imbricate aestivation
(4) Zygomorphic, epigynous with twisted aestivation
100. Lenticels are involved in
(1) Transpiration
(2) Gaseous exchange
(3) Food transport
(4) Photosynthesis
101. Age of a tree can be estimated by
(1) Its height and girth
(2) Biomass
(3) Number of annual rings
(4) Diameter of its heartwood
102. Seed coat is not thin, membranous in
(1) Maize
(2) Coconut
(3) Groundnut
(4) Gram
103. Transition state structure of the substrate formed during an enzymatic reaction is
(1) Transient but stable
(2) Permanent but unstable
(3) Transient and unstable
(4) Permanent and stable
104. A phosphoglyceride is always made up of
(1) Only a saturated fatty acid esterified to a glycerol molecule to which a phosphate group is also attached
(2) Only an unsaturated fatty acid esterified to a glycerol molecule to which a phosphate group is also attached
(3) A saturated or unsaturated fatty acid esterified to a glycerol molecule to which a phosphate group is also attached
(4) A saturated or unsaturated fatty acid esterified to a phosphate group which is also attached to a glycerol molecule
105. Pigment-containing membranous extensions in some cyanobacteria are
(1) Heterocysts
(2) Basal bodies
(3) Pneumatophores
(4) Chromatophores
106. A major site for synthesis of lipids is
(1) RER
(2) SER
(3) Symplast
(4) Nucleoplasm
107. The complex formed by a pair of synapsed homologous chromosomes is called
(1) Equatorial plate
(2) Kinetochore
(3) Bivalent
(4) Axoneme
108. The three boxes in this diagram represent the three major biosynthetic pathways in aerobic respiration. Arrows represent net reactants or products


Arrows numbered 4,8, and 12 can all be
(1) NADH
(2) ATP
(3) $\mathrm{H}_{2} \mathrm{O}$
(4) $\mathrm{FAD}^{+}$or FADH 2
109. The most abundant intracellular cation is
(1) $\mathrm{Na}^{+}$
(2) $\mathrm{Ca}^{++}$
(3) $\mathrm{H}^{+}$
(4) $\mathrm{K}^{+}$
110. During seed germination its stored food is mobilized by
(1) Ethylene
(2) Cytokinin
(3) ABA
(4) Gibberellin
111. Which of the following criteria does not pertain to facilitated transport ?
(1) Requirement of special membrane proteins
(2) High selectivity
(3) Transport saturation
(4) Uphill transport
112. The first stable product of fixation of atmospheric nitrogen in leguminous plants is
(1) $\mathrm{NO}_{2}^{-}$
(2) Ammonia
(3) $\mathrm{NO}_{3}{ }^{-}$
(4) Glutamate
113. Which of the metabolites is common to respiration mediated breakdown of fats, carbohydrates and proteins?
(1) Glucose-6-phosphate
(2) Fructose 1,6-bisphosphate
(3) Pyruvic acid
(4) Acetyl CoA
114. Which one of the following statements is correct?
(1) Hard outer layer of pollen is called intine
(2) Sporogenous tissue is haploid
(3) Endothecium produces the microspores
(4) Tapetum nourishes the developing pollen
115. Product of sexual reproduction generally generates
(1) Longer viability of seeds
(2) Prolonged dormancy
(3) New genetic combination leading to variation
(4) Large biomass
116. Meiosis takes place in
(1) Meiocyte
(2) Conidia
(3) Gemmule
(4) Megaspore
117. Advantage of cleistogamy is
(1) Higher genetic variability
(2) More vigorous offspring
(3) No dependence on pollinators
(4) Vivipary
118. Monoecious plant of Chara shows occurrence of
(1) Antheridiophore and archegoniophore on the same plant
(2) Stamen and carpel on the same plant
(3) Upper antheridium and lower oogonium on the same plant
(4) Upper oogonium and lower antheridium on the same plant
119. Perisperm differs from endosperm in
(1) Being a haploid tissue
(2) Having no reserve food
(3) Being a diploid tissue
(4) Its formation by fusion of secondary nucleus with several sperms
120. Which of the following statements is not true of two genes that show $50 \%$ recombination frequency?
(1) The genes may be on different chromosomes
(2) The genes are tightly linked
(3) The genes show independent assortment
(4) If the genes are present on the same chromosome, they undergo more than one crossovers in every meiosis
121. Variation in gene frequencies within populations can occur by chance rather than by natural selection. This is referred to as
(1) Genetic flow
(2) Genetic drift
(3) Random mating
(4) Genetic load
122. If two persons with 'AB' blood group marry and have sufficiently large number of children, these children could be classified as 'A' blood group : 'AB' blood group : 'B' blood group in 1:2:1 ratio. Modern technique of protein electrophoresis reveals presence of both 'A' and 'B' type proteins in 'AB' blood group individuals. This is an example of
(1) Codominance
(2) Incomplete dominance
(3) Partial dominance
(4) Complete dominance
123. The process by which organisms with different evolutionary history evolve similar phenotypic adaptations in response to a common environmental challenge, is called
(1) Natural selection
(2) Convergent evolution
(3) Non-random evolution
(4) Adaptive radiation
124. The tendency of population to remain in genetic equilibrium may be disturbed by
(1) Random mating
(2) Lack of migration
(3) Lack of mutations
(4) Lack of random mating
125. Which of the following Bt crops is being grown in India by the farmers?
(1) Maize
(2) Cotton
(3) Brinjal
(4) Soybean
126. A good producer of citric acid is
(1) Aspergillus
(2) Pseudomonas
(3) Clostridium
(4) Saccharomyces
127. DNA fragments generated by the restriction endonucleases in a chemical reaction can be separated by
(1) Centrifugation
(2) Polymerase chain reaction
(3) Electrophoresis
(4) Restriction mapping
128. Which of the following is not correctly matched for the organism and its cell wall degrading enzyme?
(1) Bacteria - Lysozyme
(2) Plant cells - Cellulase
(3) Algae - Methylase
(4) Fungi - Chitinase
129. The colonies of recombinant bacteria appear white in contrast to blue colonies of non-recombinant bacteria because of
(1) Non-recombinant bacteria containing betagalactosidase
(2) Insertional inactivation of alpha-galactosidase in non-recombinant bacteria
(3) Insertional inactivation of alpha-galactosidase in recombinant bacteria
(4) Inactivation of glycosidase enzyme in recombinant bacteria
130. Which of the following are likely to be present in deep sea water?
(1) Archaebacteria
(2) Eubacteria
(3) Blue-green algae
(4) Saprophytic fungi
131. Natural reservoir of phosphorus is
(1) Sea water
(2) Animal bones
(3) Rock
(4) Fossils
132. Secondary productivity is rate of formation of new organic matter by
(1) Producer
(2) Parasite
(3) Consumer
(4) Decomposer
133. Which one of the following is not used for ex situ plant conservation?
(1) Field gene banks
(2) Seed banks
(3) Shifting cultivation
(4) Botanical Gardens
134. Kyoto Protocol was endorsed at
(1) CoP - 3
(2) $\mathrm{CoP}-5$
(3) CoP - 6
(4) CoP - 4
135. Which of the following represent maximum number of species among global biodiversity?
(1) Algae
(2) Lichens
(3) Fungi
(4) Mosses and Ferns
136. Match the name of the animal (Column I) with one characteristics (Column II) and the phylum/class (column III) to which it belongs.

|  | Column I | Column II | Column III |
| ---: | :--- | :--- | :--- |
| $(1)$ | Petromyzon | Ectoparasite | Cyclostomia |
| $(2)$ | Ichthy | Terrestial | Reptila |
| $(3)$ | Limulus | Body Covered by <br> chitinous <br> exoskeleton | Pisces |
| $(4)$ | Adamsia | Radially <br> symmetrical | Porifera |

137. Which of the following are correctly matched with respect to their taxonomic classification?
(1) Flying fish, cuttlefish, silverfish, - Pisces
(2) Centipede, millipede, spider, scorpion - Insecta
(3) House fly, butterfly, tsetsefly, silverfish - Insecta
(4) Spiny anteater, sea urchin, sea cucumber - Echinodermata
138. Which group of animals belong to the same phylum?
(1) Malarial parasite, Amoeba, Mosquito
(2) Earthworm, Pinworm, Tapeworm
(3) Prawn, Scorpion, Locusta
(4) Sponge, Sea anemone, Starfish
139. One of the representatives of Phylum Arthropoda is
(1) Cuttlefish
(2) Silverfish
(3) Pufferfish
(4) Flying fish
140. The H -zone in the skeletal muscle fibre is due to
(1) The absence of myofibrils in the central portion of A-band
(2) The central gap between myosin filaments in the A-band
(3) The central gap between actin filaments extending through myosin filaments in the A-band
(4) Extension of myosin filaments in the central portion of the A-band
141. What external changes are visible after the last moult of a cockroach nymph?
(1) Mandibles become harder
(2) Anal cerci develop
(3) Both fore wings and hind wings develop
(4) Labium develops
142. The Golgi complex plays a major role
(1) In trapping the light and transforming it into chemical energy
(2) In digesting proteins and carbohydrates
(3) As energy transferring organelles
(4) In post translational modification of proteins and glycosidation of lipid
143. Which one of the following organelle in the figure correctly matches with its function?

(1) Rough endoplasmic reticulum, formation of glycoproteins
(2) Golgi apparatus, protein synthesis
(3) Golgi apparatus, formation of glycolipids
(4) Rough endoplasmic reticulum, protein synthesis
144. Macro molecule chitin is
(1) Nitrogen containing polysaccharide
(2) Phosphorus containing polysaccharide
(3) Sulphur containing polysaccharide
(4) Simple polysaccharide
145. The essential chemical components of many coenzymes are
(1) Proteins
(2) Nucleic acids
(3) Carbohydrates
(4) Vitamins
146. A stage in cell division is shown in the figure. Select the answer which gives correct identification of the stage with its characteristics.


| $(1)$ | Telophase | Nuclear envelop reforms golgi complex <br> reforms |
| :---: | :--- | :--- |
| $(2)$ | Late Anaphase | Chromosomes move away from equatorial <br> plate, golgi complex not present. |
| $(3)$ | Cytokinesis | Cell plate formed, mitochondria <br> distributed between two daughtercells. |
| $(4)$ | Telophase | Endoplasmic reticulum and nucleolus not <br> reformed yet |

147. Select the correct match of the digested products in humans given in column I with their absorption site and mechanism in column II

|  | Column I | Column II |
| :--- | :--- | :--- |
| $(1)$ | Glycine, glucose | Small intestine ,active absorption |
| $(2)$ | Fructose, Na+ | Small intestin passive absorption |
| $(3)$ | Glycerol, fatty acids | Duodenum ,move as <br> chilomicrons |
| $(4)$ | Cholesterol, maltose | Large intestine ,active absorption |

148. A pregnant female delivers a baby who suffers from stunted growth, mental retardation low intelligence quotient and abnormal skin. This is the result of
(1) Deficiency of iodine in diet
(2) Low secretion of growth hormone
(3) Cancer of the thyroid gland
(4) Over secretion of pars distalis
149. The figure shows a diagrammatic view of human respiratory system with labels $A$, B, C and D. Select the option which gives correct identification and main function and/or characteristic.

(1) A - trachea - long tube supported by complete cartilaginous rings for conducting inspired air
(2) B-pleural membrane - surround ribs on both sides to provide cushion against rubbing
(3) C-Alveoli - thin walled vascular bag like structures for exchange of gases
(4) D-Lower end of lungs - diaphragm pulls it down during inspiration
150. Figure shows schematic plan of blood circulation in humans with labels $A$ to $D$. Identify the label and give its function/s.

(1) A - Pulmonary vein - takes impure blood from body parts, $\mathrm{PO}_{2}=60 \mathrm{~mm} \mathrm{Hg}$
(2) B - Pulmonary artery - takes blood from heart to lungs, $\mathrm{PO}_{2}=90 \mathrm{~mm} \mathrm{Hg}$
(3) C - Vena Cava - takes blood from body parts to right auricle, $\mathrm{PCO}_{2}=45 \mathrm{~mm} \mathrm{Hg}$
(4) D - Dorsal aorta - takes blood from heart to body parts, $\mathrm{PO}_{2}=95 \mathrm{~mm} \mathrm{Hg}$
151. The diagram given here is the standard ECG of a normal person. The Pwave represents the

(1) Contraction of both the atria
(2) Initiation of the ventricular contraction
(3) Beginning of the systole
(4) End of systole
152. Figure shows human urinary system with structures labelled A to D. Select option which correctly identifies them and gives their characteristics and/ or functions

(1) A-Adrenal gland-located at the anterior part of kidney. Secrete Catecholamines which stimulate glycogen breakdown
(2) B-Pelvis-broad funnel shaped space inner to hilum, directly connected to loops of Henle
(3) C-Medulla - inner zone of kidney and contains complete nephrons
(4) D-Cortex - outer part of kidney and do not contain any part of nephrons
153. Select the correct statement with respect to locomotion in humans
(1) A decreased level of progesterone causes osteoporosis in old people.
(2) Accumulation of uric acid crystals in joints causes their inflammation.
(3) The vertebral column has 10 thoracic vertebrae.
(4) The joint between adjacent vertebrae is a fibrous joint.
154. The characteristics and an example of a synovial joint in humans is

|  | Characteristics | Examples |
| :--- | :--- | :--- |
| $(1)$ | Fluid cartilage <br> between two bones, <br> limited movements | Knee joints |
| $(2)$ | Fluid filled between <br> two joints, provides <br> cushion | Skull bones |
| 3$)$ | Fluid filled synovial <br> cavity between two <br> bones | Joint between <br> atlas and axis |


| $(4)$ | Lymph filled between <br> two bones, limited <br> movement | Gliding joint <br> between carpals |
| :--- | :--- | :--- |

155. A diagram showing axon terminal and synapse is given. Identify correctly at least two of A-D

(1) A - Receptor

C - Synaptic vesicles
(2) B - Synaptic connection

D $-\mathrm{K}^{+}$
(3) A - Neurotransmitter

B - Synaptic cleft
(4) C - Neurotransmitter

D $-\mathrm{Ca}^{++}$
156. Parts A, B, C and D of the human eye are shown in the diagram. Select the option which gives correct identification along with its functions/characteristics
(1) A - Retina - contains photo receptors - rods and cones.
(2) B - Blind spot - has only a few rods and cones.
(3) C - Aqueous chamber - reflects the light which does not pass through the lens.
(4) D - Choroid - its anterior part forms ciliary body.
157. Which of the following statement is correct in relation to the endocrine system?
(1) Adenohypophysis is under direct neural regulation of the hypothalamus.
(2) Organs in the body like gastrointestinal tract, heart, kidney and liver do not produce any hormones.
(3) Non - nutrient chemicals produced by the body in trace amount that act as intercellular messenger are known as hormones.
(4) Releasing and inhibitory hormones are produced by the pituitary gland.
158. Select the answer which correctly matches the endocrine gland with the hormone it secretes and its function/ deficiency symptom

|  | Endocrine <br> gland | Hormone | Function/deficiency <br> symptoms |
| :--- | :--- | :--- | :--- |
| $(1)$ | Anterior <br> pituitary | Oxytocin | Stimulates uterus contraction <br> During child birth |
| $(2)$ | Posterior <br> pituitary | Growth <br> Hormone <br> (GH) | Oversecretion stimulates <br> Abnormal growth |
| $(3)$ | Thyroid gland | Thyroxine | Lack of iodine in <br> diet results in goitre |
| $(4)$ | Corpus luteum | Testosterone | Stimulates <br> spermatogenesis |

159. What is the correct sequence of sperm formation?
(1) Spermatid, Spermatocyte, Spermatogonia, Spermatozoa
(2) Spermatogonia, Spermatocyte, Spermatozoa, Spermatid
(3) Spermatogonia, Spermatozoa, Spermatocyte, Spermatid
(4) Spermatogonia, Spermatocyte, Spermatid, Spermatozoa
160. Menstrual flow occurs due to lack of
(1) Progesterone
(2) FSH
(3) Oxytocin
(4) Vasopressin
161. Which one of the following is not the function of placenta? It
(1) Facilitates supply of oxygen and nutrients to embryo.
(2) Secretes estrogen.
(3) Facilitates removal of carbon dioxide and waste material from embryo.
(4) Secretes oxytocin during parturition.
162. One of the legal methods of birth control is
(1) Abortion by taking an appropriate medicine
(2) By abstaining from coitus from day 10 to 17 of the menstrual cycle
(3) By having coitus at the time of day break
(4) By a premature ejaculation during coitus
163. Which of the following cannot be detected in a developing foetus by amniocentesis?
(1) Klinefelter syndrome
(2) Sex of the foetus
(3) Down syndrome
(4) Jaundice
164. Artificial insemination means
(1) Transfer of sperms of a healthy donor to a test tube containing ova
(2) Transfer of sperms of husband to a test tube containing ova
(3) Artificial introduction of sperms of a healthy donor into the vagina
(4) Introduction of sperms of healthy donor directly into the ovary
165. Which Mendelian idea is depicted by a cross in which the $F_{1}$ generation resembles both the parents ?
(1) Incomplete dominance
(2) Law of dominance
(3) Inheritance of one gene
(4) Co-dominance
166. The incorrect statement with regard to Haemophilia is
(1) It is a sex-linked disease
(2) It is a recessive disease
(3) It is a dominant disease
(4) A single protein involved in the clotting of blood is affected
167. If both parents are carriers for thalassemia, which is an autosomal recessive disorder, what are the chances of pregnancy resulting in an affected child?
(1) No chance
(2) $50 \%$
(3) $25 \%$
(4) 100\%
168. The diagram shows an important concept in the genetic implication of DNA. Fill in the blanks A to C

(1) A-transcription, B-replication, C-James Watson
(2) A-translation, B-transcription, C-Erevin Chargaff
(3) A-transcription, B-translation, C-Francis Crick
(4) A-translation, B-extension, C-Rosalind Franklin
169. Which enzyme/s will be produced in a cell in which there is a nonsense mutation in the lac Y gene?
(1) $\beta$-galactosidase
(2) Lactose permease
(3) Transacetylase
(4) Lactose permease and transacetylase
170. According to Darwin, the organic evolution is due to
(1) Intraspecific competition.
(2) Interspecific competition.
(3) Competition within closely related species.
(4) Reduced feeding efficiency in one species due to the presence of interfering species.
171. The eye of octopus and eye of cat show different patterns of structure, yet they perform similar function. This is an example of
(1) Homologous organs that have evolved due to convergent evolution.
(2) Homologous organs that have evolved due to divergent evolution.
(3) Analogous organs that have evolved due to convergent evolution.
(4) Analogous organs that have evolved due to divergent evolution.
172. Infection of Ascaris usually occurs by
(1) Drinking water containing eggs of Ascaris
(2) Eating imperfectly cooked port
(3) Tse-tse fly
(4) Mosquito bite
173. The cell-mediated immunity inside the human body is carried out by
(1) T-lymphocytes
(2) B-lymphocytes
(3) Thrombocytes
(4) Erythrocytes
174. In plant breeding programmes, the entire collection(of plants/seeds) having all the diverse alleles for all genes in a given crop is called
(1) Selection of superior recombinants
(2) Cross-hybridisation among the selected parents
(3) Evaluation and selection of parents.
(4) Germplasm collection
175. During sewage treatment, biogases are produced which include
(1) Methane, hydrogensulphide, carbon dioxide
(2) Methane, oxygen, hydrogensulphide
(3) Hydrogensulphide, methane, sulphur dioxide
(4) Hydrogensulphide, nitrogen, methane
176. A biologist studied the population of rats in a barn. He found that the average natality was 250 , average mortality 240 , immigration 20 and emigration 30. The net increase in population is
(1) 10
(2) 15
(3) 05
(4) Zero
177. Which one of the following processes during decomposition is correctly described?
(1) Fragmentation - Carried out by organisms such as earthworm
(2) Humification - Leads to the accumulation of a dark coloured substance humus which undergoes microbial action at a very fast rate
(3) Catabolism - Last step in the decomposition under fully anaerobic condition
(4) Leaching - Water soluble inorganic nutrients rise to the top layers of soil
178. A sedentary sea anemone gets attached to the shell lining of hermit crab. The association is
(1) Ectoparasitism
(2) Symbiosis
(3) Commensalism
(4) Amensalism
179. Global warming can be controlled by
(1) Reducing deforestation, cutting down use of fossil fuel
(2) Reducing reforestation, increasing the use of fossil fuel
(3) Increasing deforestation, slowing down the growth of human population
(4) Increasing deforestation, reducing efficiency of energy usage
180. The Air Prevention and Control of Pollution Act came into force in
(1) 1975
(2) 1981
(3) 1985
(4) 1990a
